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CHAPTER 25 What

You'll Learn You will

trace the history of nuclear chemistry from dis-covery to

application. You will

identify types of radioactive decay and solve decay rate

problems. You will describe the reac-tions involved in nuclear

fission and fusion. You

will learn about appli-

cations of nuclear

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reactions

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Chapter 25 - Nuclear Chemistry . Isotope

Examples • Prob: -An atom of Kr has a mass of 94 AMU. How many protons & neutrons

does it have? ... 2 16

days 192 g 64 g 25% 3

24 days 224 g 32 g

12.5% 4 32 days 240 g

16 g 6.25% 5 40 days

248 g 8 g 3.125% 6 48

days 252 g 4 g

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1.5625% . Half-Life

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Chapter 25 - Nuclear Chemistry

Section 25.4 Fission and Fusion of Atomic Nuclei In your textbook, read about the process of by which electrical energy is produced in a nuclear power plant. Use the following diagram to complete the passage.

D In a nuclear power plant, energy is

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produced in reactors by fission reactions that occur in uranium-containing bars called (I)

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Chemistry

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Moonlight Interiors

Nuclear Chemistry 25.1

Nuclear Radiation 25.2

Nuclear

Transformations 25.3

Fission and Fusion 25.4

Radiation in Your Life.

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Radiation 25.2 Nuclear

Transformations 25.3

Fission and Fusion 25.4

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804 Chapter 25 Table

25.2 Decay Processes

Decay Curve for a

Radioactive Element $t_{1/2}$

2 $t_{1/2}$ $t_{1/2}$ 12 3 4

Number of half-lives 25

50 100 12.5 0

Radioisotope remaining

(%) After 1 half-life

Initial amount of

radioisotope After 2

half-lives After 3 half-

lives A nucleus may be

unstable and undergo

spontaneous

radioactive decay for

several reasons.

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**25.2 Nuclear
Transformations 25**

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Guide for Chapter 25.

Terms in this set (37)

Neutron Ejection. when

a neutron is emitted

from the nucleus. ${}^1_0\text{n}$.

Particle for Neutron

Ejection. ${}^5_2\text{He} \rightarrow {}^1_0\text{n} +$

${}^4_2\text{He}$.

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Chemistry

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Radioactivity

- Radioactivity is the process by which nuclei emit particles and rays as they break down.
- The name of the penetrating rays emitted by a radioactive source is called radiation.
- A

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radioactive isotope is
an unstable atom
which breaks down on
its own, releasing
energy and/or

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692 Chapter 16

Nuclear Chemistry 16.1

The Nucleus and

Radioactivity Our

journey into the center

of the atom begins

with a brief review. You

learned in Chapter 3

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that the protons and neutrons in each atom are found in a tiny, central nucleus that measures about $1/100,000$ the diameter of the atom itself. You also learned

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Figure 25-2 Both Pierre
and Marie Curie played
important roles in
founding the field of

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nuclear chemistry.

Marie Curie went on to show that unlike chemical reactions, radioactivity is not affected by changes in physical conditions such as temperature and pressure.

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